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Cation	Name	Comment	Anion	Name
H <sup>+</sup>	Hydrogen	Group 1A – Alkali Metal	H <sup>-</sup>	Hydride
Li <sup>+</sup>	Lithium	Group 1A – Alkali Metal	F <sup>-</sup>	Fluoride
Na <sup>+</sup>	Sodium	Group 1A – Alkali Metal	Cl <sup>-</sup>	Chloride
K <sup>+</sup>	Potassium	Group 1A – Alkali Metal	Br <sup>-</sup>	Bromide
Rb <sup>+</sup>	Rubidium	Group 1A – Alkali Metal	I <sup>-</sup>	Iodide
Cs <sup>+</sup>	Cesium	Group 1A – Alkali Metal	O <sup>-2</sup>	Oxide
			S <sup>-2</sup>	Sulfide
Be <sup>+2</sup>	Beryllium	Group 2A – Alkaline Earth Metal		
Mg <sup>+2</sup>	Magnesium	Group 2A – Alkaline Earth Metal		
Ca <sup>+2</sup>	Calcium	Group 2A – Alkaline Earth Metal		
Sr <sup>+2</sup>	Strontium	Group 2A – Alkaline Earth Metal		
Ba <sup>+2</sup>	Barium	Group 2A – Alkaline Earth Metal		
Ra <sup>+2</sup>	Radium	Group 2A – Alkaline Earth Metal		

Ag<sup>+</sup> Silver

Co <sup>2+</sup>	Cobalt	Zn <sup>+2</sup>	Zinc
Ni <sup>2+</sup>	Nickel	Mn <sup>2+</sup>	Manganese

Cr <sup>3+</sup>	Chromium	Al <sup>+3</sup>	Aluminum
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Ion	Systematic Name
Fe <sup>+2</sup> / Fe <sup>+3</sup>	Iron (II) / Iron (III)
Cu <sup>+1</sup> / Cu <sup>+2</sup>	Copper (I) / Copper (II)
Hg <sub>2</sub> <sup>+2</sup> / Hg <sup>+2</sup>	Mercury (I) / Mercury (II)

<b>1</b>	<b>Mono</b>	<b>6</b>	<b>Hexa</b>
<b>2</b>	<b>Di</b>	<b>7</b>	<b>Hepta</b>
<b>3</b>	<b>Tri</b>	<b>8</b>	<b>Octa</b>
<b>4</b>	<b>Tetra</b>	<b>9</b>	<b>Nona</b>
<b>5</b>	<b>Penta</b>	<b>10</b>	<b>Deca</b>

**PolyAtomic Ions** – several atoms bonded together – Ya gotta just memorize these:

$\text{NH}_4^+$	Ammonium	$\text{CO}_3^{-2}$	Carbonate
		$\text{HCO}_3^-$	Hydrogen Carbonate -or- Bicarbonate
$\text{NO}_2^-$	Nitrate		
$\text{NO}_3^-$	Nitrite		
		$\text{ClO}^-$	Hypo Chlorite
$\text{SO}_3^{-2}$	Sulfite	$\text{ClO}_2^-$	Chlorite
$\text{SO}_4^{-2}$	Sulfate	$\text{ClO}_3^-$	Chlorate
		$\text{ClO}_4^-$	Per Chlorate
$\text{HSO}_3^-$	Hydrogen Sulfite -or- Bisulfite		
$\text{HSO}_4^-$	Hydrogen Sulfate -or- Bisulfate		
$\text{S}_2\text{O}_3^{2-}$	Thiosulfate		
$\text{C}_2\text{H}_3\text{O}_2^-$	Acetate	$\text{C}_2\text{O}_4^{2-}$	Oxalate
$\text{OH}^-$	Hydroxide	$\text{MnO}_4^-$	Permanganate
$\text{CN}^-$	Cyanide		
$\text{PO}_4^{-3}$	Phosphate	$\text{Cr}_2\text{O}_7^{-2}$	Dichromate
$\text{HPO}_4^{-2}$	Hydrogen Phosphate	$\text{CrO}_4^{-2}$	Chromate
$\text{H}_2\text{PO}_4^-$	DiHydrogen Phosphate	$\text{O}_2^{-2}$	Peroxide

**A Trick:**

$\text{ClO}^-$	1 Oxygen	HYPO chlor ITE	Least Oxygen = HYPO
$\text{ClO}_2^-$	2 Oxygen	chlor ITE	Fewer Oxygen = ITE
$\text{ClO}_3^-$	3 Oxygen	chlor ATE	More Oxygen = ATE
$\text{ClO}_4^-$	4 Oxygen	PER chlor ATE	Most Oxygen = PER

**Common Acids:**

$\text{H}_2\text{SO}_3$	Sulfurous Acid	$\text{HF}$	Hydrofluoric Acid
$\text{H}_2\text{SO}_4$	Sulfuric Acid	$\text{HCl}$	Hydrochloric Acid
		$\text{HBr}$	HydroBromic Acid
$\text{HNO}_2$	Nitrous Acid	$\text{HI}$	HyrdoIodic Acid
$\text{HNO}_3$	Nitric Acid		
		$\text{HClO}$	Hypo Chlorous acid
$\text{H}_3\text{PO}_4$	Phosphoric Acid	$\text{HClO}_2$	Chlorous Acid
$\text{HC}_2\text{H}_3\text{O}_2$	Acetic Acid [ diluted is vinegar ]	$\text{HClO}_3$	Chloric Acid
		$\text{HClO}_4$	Per Chloric Acid
$\text{H}_2\text{CO}_3$	Carbonic Acid		