

Chapter 6: Chemical Reactions, An Introduction

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Example Reactions. Show Balanced Equations Only

Hydrochloric Acid and Cobalt (II) Nitrate

Sodium Dichromate and Lead (II) Nitrate

Methane and Oxygen [Burn Methane]

Hydrofluoric Acid and silicon Dioxide

Hydrochloric Acid and Sodium Hydroxide

Sulfuric Acid and Potassium Hydroxide

Phosphoric Acid and Magnesium Hydroxide

Sulfuric Acid and Aluminum Hydroxide

Silver Nitrate and Sodium Chloride

Silver Nitrate and Barium Chloride

Copper (I) Sulfite and Phosphoric Acid

Copper (II) Hypochlorite and Nitric Acid

Iron (II) Chlorate and Potassium Hydroxide

Iron (III) Carbonate and Sulfuric Acid

Cobalt (II) Nitrate and Sodium Iodide

Cobalt (III) Sulfate and Barium Carbonate

Tin (II) Hydroxide and Magnesium Phosphate

Tin (IV) Bisulfate and Cobalt (II) Chromate

Lead (II) Acetate and Iron (III) Chromate

Lead (IV) Permanganate and Iron (III) Cyanide

Ammonium Dichromate and Iron (III) Permanganate

Lithium Chromate and Silver Phosphate

Zinc Bicarbonate and Lithium Hydrogen Phosphate

Magnesium Hydroxide and Hydrochloric Acid

Silver Nitrate and Copper (II) Chloride

Iron (II) Phosphate and Aluminum Hydroxide

Cobalt (III) Sulfate and Potassium Chloride

Ammonium Perchlorate and Sodium Hydroxide

Calcium Hydroxide and Hydrochloric Acid

Sulfuric Acid and Iron (III) Hydroxide

Copper (II) Carbonate and Phosphoric Acid

Potassium Dichromate and Lithium Chloride

Balance the following and say if the reaction will go to completion and why?

If it does go to completion, show complete ionic and net ionic reactions

Sodium Chloride and Magnesium Bromide

Potassium Hydroxide and Iron (III) Chloride

Barium Nitrate and Potassium Phosphate

Sodium Sulfate and Potassium Chloride

Nickel (II) Nitrate and Potassium Carbonate

Sodium Sulfide and Copper (II) Nitrate

Ammonium Chloride and Lead (II) Nitrate

Nickel Chloride and Lithium Carbonate

Lead (II) Nitrate and Potassium Iodide

Barium Chloride and Sodium Sulfate

Copper (II) Chloride and Sodium Carbonate

Hydrochloric Acid and Lithium Hydroxide

Sulfuric Acid and Calcium Hydroxide

Zinc and Hydrochloric Acid

Aluminum oxidizes with Oxygen to form oxide

Sodium reacts with Oxygen to form Sodium Oxide

Hydrochloric Acid and Sodium Carbonate

Hydrogen and Oxygen burn to form water

Water decomposes to Hydrogen and Oxygen ↑

Mercury (II) Oxide decompose to Mercury and Oxygen

Barium Chloride and Sodium Sulfate

Iron (III) Chloride and Phosphoric Acid

Barium Nitrate and Sodium Sulfate