

Chem 1025 Chapter 6 Equations

- Write formulae for the Reactants and Products
- Balance the equation
- Will the reaction go to completion

Hydrochloric Acid and Cobalt III Nitrate -> Blue
 $\text{HCl} + \text{Cobalt III Nitrate} \rightarrow \text{Blue}$

Sodium Dichromate + Lead Nitrate -> Solid

Ammonium Dichromate -> Chromium III Oxide + Nitrogen + Water
 $(\text{NH}_4)_2\text{Cr}_2\text{O}_7 \rightarrow \text{Cr}_2\text{O}_3 + \text{N}_2 + \text{HOH}$

Ammonia + Oxygen -> $\text{NO} + \text{HOH}$

Calcium + Water -> Bubbles

Methane + Oxygen -> Flame

Zinc + HCl ->

Magnesium Metal and Water ->
 $\text{Mg} + \text{HOH} \rightarrow \text{Mg}(\text{OH})_2 + \text{H}_2$

Iron III Oxide + Nitric Acid -> Iron III Nitrate + Water
 $\text{Fe}_2\text{O}_3 + 6 \text{HNO}_3 \rightarrow 2 \text{Fe}(\text{NO}_3)_3 + 3 \text{HOH}$

Hydrogen Sulfide gas + Lead II Nitrate -> Lead II Sulfide + Nitric Acid
 $\text{H}_2\text{S} + \text{Pb}(\text{NO}_3)_2 \rightarrow \text{PbS} + 2 \text{HNO}_3$

Iron III Chloride + Potassium Hydroxide
 $\text{FeCl}_3 + 3 \text{KOH} \rightarrow \text{Fe}(\text{OH})_3 + 3 \text{KCl}$

Lead II Acetate + Potassium Iodide
 $\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_2 + 2 \text{KI} \rightarrow \text{PbI}_2 + 2 \text{K C}_2\text{H}_3\text{O}_2$

Lithium Oxide + Water ->
 $\text{Li}_2\text{O} + \text{HOH} \rightarrow 2 \text{LiOH}$

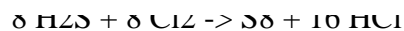
Antimony Metal and Chlorine Gas ->
 $\text{Sb} + 3 \text{Cl}_2 \rightarrow 2 \text{SbCl}_3$

Iron II Sulfide + Hydrochloric Acid ->
 $\text{FeS} + 2 \text{HCl} \rightarrow \text{FeCl}_2 + \text{H}_2\text{S}$

Aluminum Metal and Copper II Oxide ->
 $\text{Al} + 3 \text{CuO} \rightarrow \text{Al}_2\text{O}_3 + 3 \text{Cu}$

Ammonium Chloride + Potassium Hydroxide
 $\text{NH}_4\text{Cl} + \text{KOH} \rightarrow \text{NH}_3 + \text{HOH} + \text{KCl}$

Hydrogen Sulfide + Chlorine Gas ->



Lithium Hydroxide + Carbon Dioxide
 $2 \text{LiOH} + \text{CO}_2 \rightarrow \text{Li}_2\text{CO}_3 + \text{HOH}$

Potassium Carbonate + Nitric Acid ->
 $\text{K}_2\text{CO}_3 + 2 \text{HNO}_3 \rightarrow 2 \text{KNO}_3 + \text{HOH} + \text{CO}_2$

Sodium Chloride + Sulfuric Acid ->
 $2 \text{NaCl} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + 2 \text{HCl}$

Sodium Sulfate + Calcium Chloride ->
 $\text{Na}_2\text{SO}_4 + \text{CaCl}_2 \rightarrow \text{CaSO}_4 + 2 \text{NaCl}$

Calcium Hydroxide + Hydrochloric Acid ->
 $\text{Ca}(\text{OH})_2 + 2 \text{HCl} \rightarrow \text{CaCl}_2 + 2 \text{HOH}$

Sodium Hydroxide + Phosphoric Acid ->
 $3 \text{NaOH} + \text{H}_3\text{PO}_4 \rightarrow \text{Na}_3\text{PO}_4 + 3 \text{HOH}$

Sodium Bicarbonate -> Decomposition with heat
 $2 \text{NaHCO}_3 \rightarrow \text{Na}_2\text{CO}_3 + \text{CO}_2 + \text{HOH}$

Barium Nitrate + Sodium Chromate ->
 $\text{Ba}(\text{NO}_3)_2 + \text{Na}_2\text{CrO}_4 \rightarrow \text{BaCrO}_4(\text{s}) + 2 \text{NaNO}_3$

Lead II Chloride + Potassium Sulfate ->
 $\text{PbCl}_2 + \text{K}_2\text{SO}_4 \rightarrow \text{PbSO}_4 + 2 \text{KCl}$

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Iron III Chloride + Potassium Hydroxide

Lead II Acetate + Potassium Iodide

Lithium Oxide + Water ->

Antimony Metal and Chlorine Gas ->

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Aluminum Metal and Copper II Oxide ->

Ammonium Chloride + Potassium Hydroxide

Hydrogen Sulfide + Chlorine Gas ->

Lithium Hydroxide + Carbon Dioxide

Potassium Carbonate + Nitric Acid ->

Sodium Chloride + Sulfuric Acid ->

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Calcium Hydroxide + Hydrochloric Acid ->

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